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CLASS – X

HOLIDAY HOMEWORK CHEMISTRY

- (1) When a green iron salt is heated strongly its colour finally changes to brown and odour of burning sulphur is given out:
- Name the iron salt.
 - Name the type of reactions that takes place during the heating of iron salt.
 - Write a chemical equation for the reaction involved.
- (2) A colourless lead salt, when heated, produces a yellow residue and brown fumes:
- Name the lead salt
 - Name the brown fumes
 - Write the chemical equations
- (3) (a) What happens when an aqueous solution of sodium sulphate reacts with an aqueous solution of barium chloride.
- (b) Write the balanced chemical equation for the reaction.
- (c) State the physical conditions of reactants in which the reactions will not takes place.
- (d) Name the type of chemical reaction.
- (e) Give an example of another reaction which is of the same type as the above reaction.
- (4) (a) Define a combination reaction.
- (b) Give an example of a combination reaction which is also exothermic.
- (c) Give an example of a combination reaction which is also endothermic.